

Module Handbook

Module Name:	General Chemistry
Module Level:	Bachelor
Abbreviation, if applicable:	Lecture KID105 Practical Work KID106
Sub-heading, if applicable:	
Courses included in the module, if applicable:	
Semester/term:	1 / First year
Module coordinator(s):	Prof. Dr. Noor Erma NS, MS., Apt.
Lecturer(s):	Prof. Dr. Noor Erma NS, MS., Apt. Prof. Dr. Amirudin Prawita, Apt. Dr. Bambang Tri P., MS., Apt. Dra. Asri Darmawati, MS., Apt. Drs. A. Toto P., MSi., Apt. Dr. Juni Ekowati, MSi., Apt. Dr. Riesta Primaharinastiti, S.Si., M.Si. Apt. Febri Annuryanti, S.Farm., M.Sc., Apt. Kholis Amalia N, S.Farm., M.Sc., Apt.
Language:	Bahasa Indonesia
Classification within the curriculum:	Compulsory Course/ Elective Studies
Teaching format/class hours per week during the semester:	Lecture 100 minutes lectures, 13 lecture classes/semester Practical Work 100 minutes practical work classes, 13 practical work classes /semester
Workload:	Lecture Total 22 hours a semester Practical Work Total 22 hours a semester
Credit Points:	Lecture 2 Practical Work 1
Requirements:	
Learning goal/competencies:	Knowledge <ul style="list-style-type: none"> - To understand the concept substance; and basic concepts and principles in basic chemistry. Skills <ul style="list-style-type: none"> - Critical thinking, comprehensive and valid operating in scientific-academic, active learning for accessing information, make scientific decision-academic. Competence <ul style="list-style-type: none"> - To understand and able to apply the basic theories of chemistry, so it can be used to support courses / advanced practicum.
Content:	Lecture Atoms and elements; molecules and compounds; the substance of system, the balance of ions in solution;

	<p>electrochemistry</p> <p>Practical Work</p> <p>Atoms and elements; molecules and compounds; the substance of system, the balance of ions in solution; electrochemistry</p>
Study/exam achievements:	<p>Lecture</p> <p>Student are considered to be competent and pass if at least get 50% of maximum mark of the exams based learning.</p> <p>Final score is calculated as follow :</p> <p>50% Exam I + 50% Exam II</p> <p>Final index is defined as follow :</p> <p>A : ≥ 75</p> <p>AB : 70 – 74,9</p> <p>B : 65 – 69,9</p> <p>BC : 60 – 64,9</p> <p>C : 55 – 59,9</p> <p>D : 40 – 54,9</p> <p>E : <40</p> <p>Practical Work</p> <p>Student are considered to be competent and pass if at least get 50% of maximum mark of the exams based learning.</p> <p>Final score is calculated as follow :</p> <p>30% Pretest + 20% Report + 50% Final Exam</p> <p>Final index is defined as follow :</p> <p>A : ≥ 75</p> <p>AB : 70 – 74,9</p> <p>B : 65 – 69,9</p> <p>BC : 60 – 64,9</p> <p>C : 55 – 59,9</p> <p>D : 40 – 54,9</p> <p>E : <40</p>
Forms of Media:	Slides and LCD Projector, whiteboards, internet.
Literature:	<ol style="list-style-type: none"> 1. Bishop <i>et.al.</i>, 1992. Experiments in General Chemistry, 2nd Ed. Harcourt BraceCollegePublishers, New York. 2. Brady JE, Holum JR., 1994. Chemistry, the Study of Matter and its Changes, John Wiley & Sons Inc, New York. 3. Christian GD., 1994. Analytical Chemistry, John Wiley & Sons, New York. 4. Hein M <i>et al.</i>, 1992. College Chemistry in the Laboratory, 5th Ed. Broke/Cole Publishing, California.
Notes:	The course is more concept to chemistry based as compared to basic chemistry.

